

Balancing Redox $\frac{1}{2}$ Reaction (Basic)

Balancing Basic

- Balancing in basic solution is just like acidic with one additional step
- When balancing in acidic you will have H^+ at the end
- In basic you don't want H^+ , you want OH^-

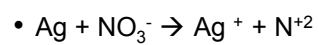
To change to basic

- To change to basic, simply add an equivalent number of OH^- to each side and simplify

Example

- Balancing the following by using the $\frac{1}{2}$ reaction method in a basic solution
- $AsO_4^{-3} + Zn \rightarrow H_3As + Zn^{+2}$

Another Example



One More

