

AP Chemistry Ion List

1+	Ammonium	NH_4^+	Potassium	K^+
	Copper (I) or Cuprous	Cu^+	Silver	Ag^+
	Hydrogen	H^+	Sodium	Na^+
	Hydronium	H_3O^+	Cesium	Cs^+
	Lithium	Li^+		

2+	Barium	Ba^{2+}	Magnesium	Mg^{2+}
	Cadmium	Cd^{2+}	Manganese (II)	Mn^{2+}
	Calcium	Ca^{2+}	Mercury (I) or Mercurous	Hg_2^{2+}
	Chromium (II)	Cr^{2+}	Mercury (II) or Mercuric	Hg^{2+}
	Cobalt (II)	Co^{2+}	Nickel (II)	Ni^{2+}
	Copper (II) or Cupric	Cu^{2+}	Strontium	Sr^{2+}
	Iron (II) or Ferrous	Fe^{2+}	Tin (II) or Stannous	Sn^{2+}
	Lead (II) or Plumbous	Pb^{2+}	Zinc	Zn^{2+}

3+	Aluminum	Al^{3+}	Chromium (III)	Cr^{3+}
	Arsenic (III)	As^{3+}	Cobalt (III)	Co^{3+}
	Bismuth (III)	Bi^{3+}	Iron (III) or Ferric	Fe^{3+}

4+	Lead (IV) or Plumbic	Pb^{4+}	Tin (IV) or Stannic	Sn^{4+}
	Manganese (IV)	Mn^{4+}		

5+	Arsenic (V)	As^{5+}	Bismuth (V)	Bi^{5+}
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1 -	Bromide	Br^-	Hydrogen Carbonate or Bicarbonate	HCO_3^-
	Perbromate	BrO_4^-	Hydrogen Oxalate or Binoxalate	HC_2O_4^-
	Bromate	BrO_3^-	Hydrogen Sulfate or Bisulfate	HSO_4^-
	Bromite	BrO_2^-	Hydrogen Sulfite or Bisulfite	HSO_3^-
	Hypobromite	BrO^-	Hydrogen Sulfide or Bisulfide	HS^-
	Chloride	Cl^-	Hydroxide	OH^-
	Perchlorate	ClO_4^-	Iodide	I^-
	Chlorate	ClO_3^-	Periodate	IO_4^-
	Chlorite	ClO_2^-	Iodate	IO_3^-
	Hypochlorite	ClO^-	Iodite	IO_2^-
	Cyanide	CN^-	Hypoiodite	IO^-
	Dihydrogen Phosphate	H_2PO_4^-	Permanganate	MnO_4^-
	Ethanoate or Acetate	$\text{C}_2\text{H}_3\text{O}_2^-$	Nitrate	NO_3^-
	Fluoride	F^-	Nitrite	NO_2^-
Hydride	H^-	Thiocyanate	SCN^-	

2 -	Carbonate	CO_3^{2-}	Peroxide	O_2^{2-}
	Chromate	CrO_4^{2-}	Sulfite	SO_3^{2-}
	Dichromate	$\text{Cr}_2\text{O}_7^{2-}$	Sulfide	S^{2-}
	Hydrogen Phosphate	HPO_4^{2-}	Sulfate	SO_4^{2-}
	Oxalate	$\text{C}_2\text{O}_4^{2-}$	Thiosulfate	$\text{S}_2\text{O}_3^{2-}$
	Oxide	O^{2-}		

3 -	Nitride	N^{3-}	Phosphate	PO_4^{3-}
	Phosphide	P^{3-}	Phosphite	PO_3^{3-}

Acid Rules	H + _____ - ide	=	hydro <u>ROOT</u> - ic	Acid
	H + _____ - ate	=	<u>ROOT</u> - ic	Acid
	H + _____ - ite	=	<u>ROOT</u> - ous	Acid